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138 Chapter 5 • Electrons in Atoms Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the

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Equation Test on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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occupies the lowest energy
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Q. The Quantum Mechanical Model of the Atom describes the electron's probable location around the nucleus in a 3-D cloud called a(n) _____.

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Atoms Test certain orbits located around the nucleus. Every electron has a fixed energy in certain energy levels: they are said to be quantized. Electrons farther away from the nucleus have higher energy than electrons

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Atoms to the nucleus.

Energy levels are closer together the farther away from the nucleus.

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in order of their atomic numbers so that elements with similar properties fall in the same column, or group. ... Electrons are added as well, but ...

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The orbital diagram for a ground-state nitrogen atom is $1s^2 2s^2 2p^3$. The number of orbitals in a d subshell is 5. The maximum number of electrons that can

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occupy an energy level described by the principal quantum number, n , is A: $2n^2$ 7. A ground-state atom of manganese has _____ unpaired electrons and is _____. A: 5, paramagnetic 8.

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156: 106. Answer. The atomic
mass of chlorine is very far
from a whole because a

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weighted average of atomic masses of all of its isotopes is computed in determining its atomic mass.
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electrons exhibit properties
of both particles and waves.

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d. the chemical properties of elements can be grouped according to periodicity but physical properties cannot.

_____ 28. Elements in a group or column in the periodic table can be expected to have similar. ... Chemistry

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In Chemistry, resonance is a
way of describing
delocalized electrons in an

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atom or molecule that cannot be represented with a single Lewis dot structure. ...
Determine the total # of valence electrons in a molecule: N 5. 30 18. Ve ...

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